Tube No.201SACHLORIDE ION

Solution 🔿

1. PERFORMANCE

| 1) Sampling method | : Immersion method | |
|--------------------------|---------------------------------|--|
| | (Refer to Page 17) | |
| 2) Measuring range | : 10-2,000 ppm | |
| 3) Sampling time | : 1.5 minutes | |
| 4) Sample volume | : over 5 m ℓ | |
| 5) Detectable limit | : 5 ppm | |
| 6) Shelf life | : 3 years | |
| 7) Operating temperature | : 5 ∼ 80 °C | |
| 8) Operating PH | : 3.5-13 | |
| 9) Reading | : Direct reading from the scale | |
| 10) Colour change | : Brown→White | |

2. RELATIVE STANDARD DEVIATION

RSD-low: 10% RSD-mid.: 10% RSD-high: 10%

3. CHEMICAL REACTION

By reacting with Silver chromate, Silver chloride is produced. $CI^- + Ag_2CrO_4 \rightarrow AgCI$

4. CALIBRATION OF THE TUBE

SODIUM CHLORIDE STANDARD SOLUTION METHOD

5. INTERFERENCE AND CROSS SENSITIVITY

| Substance | Interference | Coexistence |
|--------------|--------------------------|--|
| Bromide ion | | Higher readings are given. |
| lodide ion | | " |
| Cyanide ion | | // |
| Sulphide ion | Brown stain is produced. | The bottom of discoloured layer is changed to Brown and higher readings are given. |

6. SAMPLING METHOD

(Immersion method)

- 1) Cut both ends of a fresh detector tube with a file.
- 2) Immerse the filled end of the tube with white end plug into the prepared sample solution.

Capillary action will occur immediately and the sample solution rise through the reagent. If Chloride ion is existed in the sample solution, a discolouration will be occurred in the detecting reagent layer from its inlet and the discoloured layer shall be given according to the concentration of Chloride ion.

SAMPLE SOLUTION WHITE END PLUG